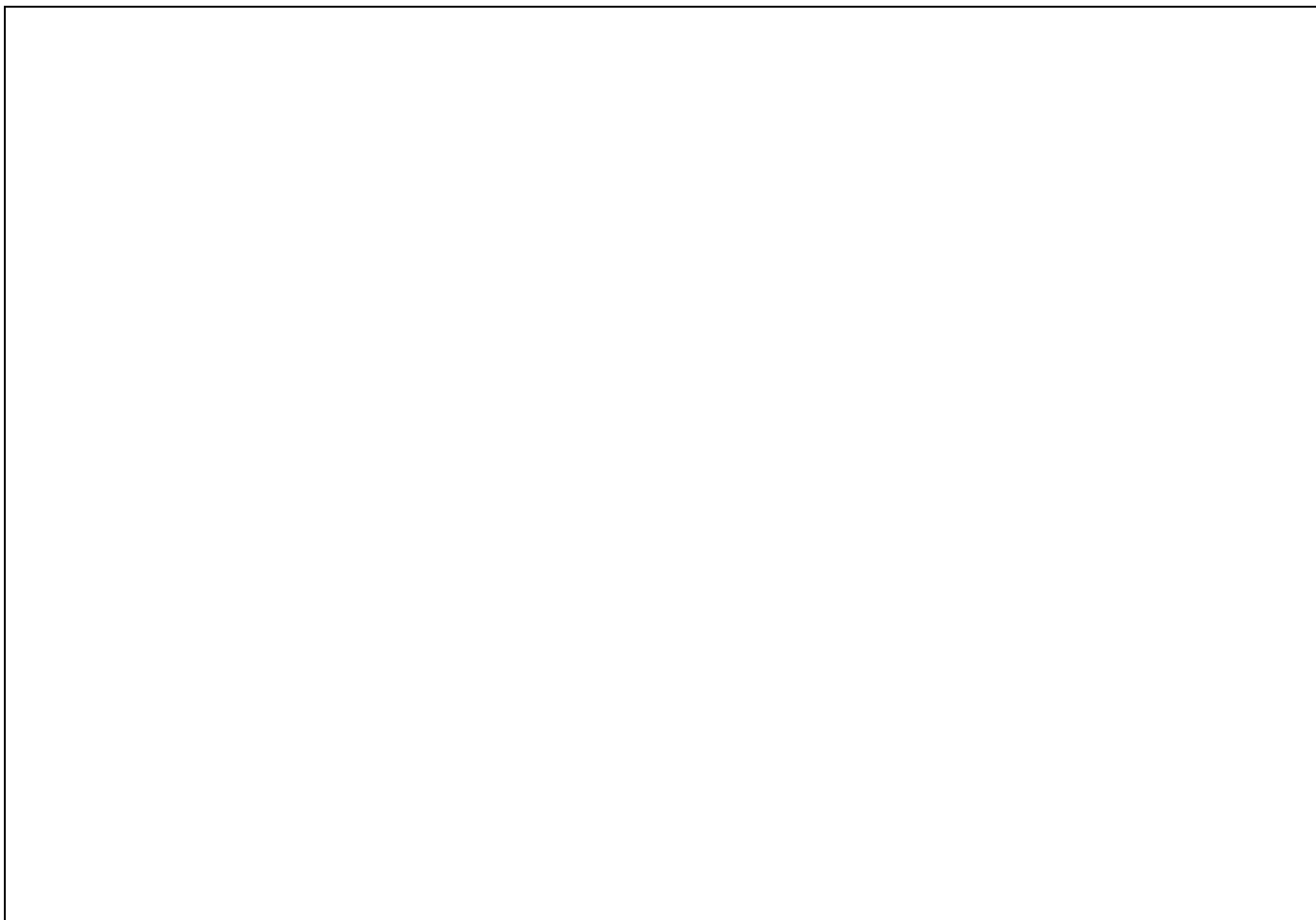


## Electrolysis Of Copper (II) Bromide Solution

1. Draw a diagram summarising the electrolysis of aqueous copper (II) bromide. In your diagram you must include the cell, electrodes, wires, ions, electrolyte and at which electrodes the products are formed



2. What is formed at the cathode?

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3. What is formed at the anode?

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4. Why is it easier to carry out electrolysis on an aqueous solution compared to a molten liquid?

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5. Why can solid ionic compounds not be electrolysed?

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(H) 6. Write a half equation for  $\text{Br}^-$  forming  $\text{Br}_2$

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(H) 7. Write a half equation for  $\text{Cu}^{2+}$  forming Cu

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8. What is left in the solution?

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